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Modern Chemistry 140 Chapter Test Name Class Date Chapter Test B, continued \_\_\_\_ 7. An acid-base titration determines the solution volumes that are a. chemically equivalent. b. of equal molarity. c. of equal mass. d. of equal molality. \_\_\_\_ 8. In acidic solutions, an indicator that is a weak acid, HIn, is primarily in the form a. In . b. In ...

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### Holt Modern Chemistry Assessment Answer Key

of solution. Molality is the number of moles of solute per kilogram of solvent. 27. Solutions are homogeneous mixtures. Suspensions and colloids are heteroge-neous mixtures. Solutions have the smallest particle size, followed by the size of colloid particles and suspen-sion particles. Solutions and colloids do not settle out on standing, but ...

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Assessment Solutions Class Date Section Quiz: Concentration of Solutions In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. 1. When preparing 500. ml, of a 1.35 M aqueous solution of NaCl, what should you do after adding the correct amount of solute to a large beaker? a.

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b. color of the solution. c. concentration of the solution. d. voltage difference between two electrodes in the meter's probe. \_\_\_\_15. When you conduct an acid-base titration, a. the pH of the solution must go up. b. the pH of the solution must go down. c. the pH of the solution must be 7.0 at the end point. d. the equivalence point must be ...

### Assessment Chapter Test A - Baumapedia

In acidic solutions, the H<sup>30</sup> ions in solution drive the equation toward the nonionized form. HIn is present in largely nonionized form in acidic solutions, and In ions are present in largely ionized form in basic solutions. HIn is a different color than the In ion. 29. acidic 30. neutral 31. basic 32. basic 33. basic 34. [H<sup>30</sup>] 1 10<sup>4</sup> M; [OH ...

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