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## **Holt Chemistry Stoichiometry Chapter Summary**

The stoichiometry of a reaction describes the relative amounts of reactants and products in a balanced chemical equation. A stoichiometric quantity of a reactant is the amount necessary to react completely with the other reactant (s). If a reactant remains unconsumed after complete reaction has occurred, it is in excess.

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relationship between two or more substances d.... the substance that controls the quantity of product that can f.... the substance that is not used up completely in a reaction.

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STOICHIOMETRY : Stoichiometry is the calculations of the quantities of reactants and products involved in a chemical reaction. Following methods can be used for solving problems.

(a) Mole Method (For Balance reaction) (b) POAC method }

Balancing not required but common sense ----- use it with slight

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care.

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Stoichiometry is at the heart of the production of many things you use in your daily life. Soap, tires, fertilizer, gasoline, deodorant, and chocolate bars are just a few commodities you use that are chemically engineered, or produced through chemical reactions. Chemically engineered commodities all rely on stoichiometry for their production.

## **Introduction to Stoichiometry: Overview | SparkNotes**

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Chemistry I Honors class. Students of high school and college general chemistry may find them useful as a supplement to their own class notes or as a review. Teachers, please feel free to use and modify them for your own classes.

## **Mrs. J's Chemistry Page - Lecture Notes**

We hope your visit has been a productive one. If you're having any problems, or would like to give some feedback, we'd love to hear from you. For general help, questions, and suggestions, try our dedicated support forums. If you need to contact the CourseNotes.Org web experience team, please use our contact form.

## **Unit 4 - Conservation of Mass and Stoichiometry | CourseNotes**

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