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Chemical Exergy  
T K P 101 325 Kpa

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## **Appendix 1 Standard Chemical Exergy**

1 APPENDIX 1.

STANDARD CHEMICAL  
EXERGY ( $T_n = 298.15$   
K,  $p_n = 101.325$  kPa)

Source: Szargut, J.,  
Egzergia. Poradnik  
obliczania i stosowania,  
Wydawnictwo  
Politechniki Śląskiej,  
Gliwice 2007.

Substance State

Molecular mass

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Chemical Exergy  
Enthalpy of  
devaluation Standard  
chemical exergy  $M$ ,  
kg/kmol  $D_o$ , kJ/mol  $e_o$   
 $x, ch$ , kJ/mol  $Al$  s  
26.9815 930.9 795.7

## **APPENDIX 1. STANDARD CHEMICAL EXERGY (T K, p 101.325 kPa)**

Chemical exergy 1 .  
CHEMICAL EXERGY.  
Table 1. Molar exergy,  
 $\psi_i$ , of pure species  
relative to a reference  
atmosphere at p

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Chemical Exergy  
T<sub>0</sub>=100 kPa, T<sub>0</sub>=25 °C,  
and 60% RH;  $\psi_i = \mu_i(T_0, p_{0,i}) - \mu_i^0(T_0, p_{0,i}, x_i^0)$ . Substance Formula  
(state)  $\psi$  Molar fraction  
in ref. atm.  $x_i$  Molar  
exergy  $\cdot i$  [kJ/mol]  
Nitrogen N<sub>2</sub> (g)  
0.7651 0.66  
Oxygen O<sub>2</sub> (g)  
0.2062 3.9  
Water H<sub>2</sub>O (l)

## **CHEMICAL EXERGY - UPM**

APPENDIX 1.  
STANDARD CHEMICAL  
EXERGY in

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Chemical Exergy  
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Thermodynamics and  
the Destruction of

Resources by Bhavik  
Bakshi, Timothy  
Gutowski and Dusan  
Sekulic, Cambridge  
University Press, 2011 .  
Materials Production.

Ayres et al, An  
Application of Exergy  
Accounting to Five  
Basic Metal Industries,  
2006. Ellingham  
Diagrams.

**2.83/2.813 Reading  
List - Massachusetts**



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## Chemical Exergy Institute of ...

Merely said, the appendix 1 standard chemical exergy t k p 101 325 kpa is universally compatible in the same way as any devices to read. It may seem overwhelming when you think about how to find and download free ebooks, but it's actually very simple.

## Appendix 1 Standard Chemical Exergy T K

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## **P 101 325 Kpa**

Starting from eq. , the standard molar chemical exergy of the compound "X" at the standard condition (T = 298.15 K and P = 1 atm) is defined as in

$$\begin{aligned} \epsilon^{\circ} = & \Delta H_f^{\circ} - T \Delta S_f^{\circ} + a \epsilon^{\circ} \text{C} + b 2 \epsilon^{\circ} \text{H}_2 \\ & + c 2 \epsilon^{\circ} \text{N}_2 + d 2 \epsilon^{\circ} \text{O}_2 + e \epsilon^{\circ} \text{S} + f 2 \epsilon^{\circ} \text{F}_2 \\ & + g 2 \epsilon^{\circ} \text{Cl}_2 + h 2 \epsilon^{\circ} \text{Br}_2 + i 2 \epsilon^{\circ} \text{I}_2 + j \epsilon^{\circ} \text{S}_i \end{aligned}$$

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Chemical Exergy  
**Standard molar Kpa  
chemical exergy: A  
new accurate model**

...

The standard exergy of a substance can be evaluated by considering an idealized reaction of the substance with other substances (generally reference substances) of known chemical exergies 16. The known chemical exergies can be obtain

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Chemical Exergy  
TKPa  
from the table of  
standard chemical  
exergy based on  
Szargut's R.E model  
(Model II), as shown in  
Table 1.

## **Estimating the specific chemical exergy of municipal solid ...**

Answers for 1 & 2 !S=S  
2 "S 1 = 1 T o #\$(E 2  
"E 1)"(B 2 "B 1)%&= Q  
L T L " Q H T H B lost  
=Q H "Q L +T o!S"W  
Consider the process in

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two stages; 1) you transfer heat in, and 2) You transfer heat and work out. Use the result from Carnot to Show that the change in entropy is zero. This leads to the result that The exergy lost is also zero.

## **Introduction to the “Exergy” Concept**

Definition #1 of Exergy  
“Exergy is the amount of work obtainable when some matter is

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brought to a state of thermodynamic equilibrium with the common components of the natural surroundings by means of reversible processes...” [Szargut et al 1988].

## **Introduction to the “Exergy” Concept**

Appendix 1. Standard chemical exergy. 2007.

[30] Querol E, Gonzalez-Regueral B, and Perez-Benedito JL, Practical

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ap- proach to exergy  
and thermo-economic  
analysis of industrial

processes. 2013;

Springer. [31] Cullen  
JM, Allwood JM,

Borgstein EH. Reducing  
Energy Demand : What

Are the Practical  
Limits? Environ Sci  
Technol.

2011;45:17118. [32]

**Unlocking Plant-  
level Resource  
Efficiency Options: A**

...

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Fe (c = 1) 376.4

kJ/mole reduction Fe

2O<sub>3</sub> (c=1) 16.5 kJ/mole

extraction Fe 2O<sub>3</sub>(c =

1.3 x 10<sup>-3</sup>) 0 kJ/mole

(ground)

## **Introduction to the concept of available work and exergy**

Standard chemical exergy and standard entropies of various compounds. Substance e 0 (kJ/mol) s 0 (kJ/mol K) ... (Appendix 1).

These data cover six



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Chemical Exergy  
categories of .

TKPa for MSW Kpa  
fractions, namely ...

## **(PDF) Estimating the specific chemical exergy of municipal**

...

Standard chemical  
exergy values, in units  
of kJ/kmol, are based  
on a standard exergy  
reference environment  
with  $T_0$  and  $P_0$  such as  
298.15 K (536.67 R)  
and 1 atm,  
respectively, and

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**Chemical Exergy**  
TK 101 225 kPa  
consisting of a set of  
reference substances  
with standard  
concentrations of  
gaseous, liquid and  
solid components.

## **Chemical Exergy - an overview |**

### **ScienceDirect Topics**

Standard chemical  
exergy of a substance.  
The chemical exergy of  
a substance is equal to  
the maximum amount  
of work that can be  
obtained from the

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Chemical Exergy  
TKP 101 2017a

substance by taking it to chemical equilibrium with the reference environment .

## **Estimating the specific chemical exergy of municipal solid ...**

subatomic particles prevents a standard exergy for each nuclide in a form similar to chemical species standard exergy.

Nuclides have specific bound energies [7]

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Chemical Exergy  
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following a curve rising on either side of the atoms in the iron family of the periodic table as shown in Fig. 1. Bound energy is proportional to the mass difference between a

## **Quantifying global exergy resources - thermophysics.ru**

1.013 bar. Standard chemical exergy values of the main substances were obtained using

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the model of Szargut [43]. ... simulation of the plant can be found in the Appendix A of the paper.

## **(PDF) Exergy and Economic Evaluation of a Hybrid Power ...**

The specific chemical exergy can also be calculated using Eq. (Mehrpooya and Mousavi, 2018): (27)  $e_{ch} = \sum x_i e_{i0} + G - \sum x_i G_i$  where  $x_i$ ,  $e_{i0}$ , and  $G_i$  are the mole

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fraction, standard  
chemical exergy of  
component  $i$  for an  
ideal mixture, and the  
Gibbs free energy.  
Exergy balance around  
the device, under  
steady state conditions  
can be ...

## **Batch-to-continuous process design and economic, energy ...**

of universal and  
comprehensive  
formulations of the  
chemical exergy

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reference environment,  
for the purpose of  
better understanding  
how exergy can inform  
decision-making. Four  
related insights  
emerged from the  
analysis, notably: (1)  
standard and universal  
chemical exergy  
reference  
environments  
necessarily encounter  
internal inconsistencies  
and even

**The Characteristics**

*Page 23/27*

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## Chemical Exergy Reference 325 Kpa Environment ...

The exergy loss ratio is calculated as : (11)  $E_{loss} = T_0 \int \gamma dx_e$  where  $e_f$  is the initial chemical exergy carried by the mixture, defined as the product between the standard chemical exergy per unit mass of the unburned mixture (see Appendix B for details), the density of the unburned mixture and



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Chemical Exergy  
the laminar flame  
speed: (12) e f...  
TKP (10) 325 Kpa

## **Second law thermodynamic analysis of syngas premixed ...**

@article{osti\_6157620,  
title = {Exergy  
analysis of thermal,  
chemical, and  
metallurgical  
processes}, author =  
{Szargut, J and Morris,  
D R and Steward, F R},  
abstractNote = {This  
important new text

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demonstrates exergy applications in ecology, economics, and industry. In fact, it constitutes the very first presentation of the vast possibilities of important ecological applications of exergy analysis.

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ecf8427e.

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